Element	Key Oxidation Numbers	
oxygen	**-2 (oxide)**	**-1 (peroxide)**
	1.) Variable subscript for oxygen 2.) Any metal and some nonmetals with oxygen: M_xO_y	1.) A subscript of 2 for oxygen 2.) Mostly Group IA and IIA metals with oxygen: M_xO_2
	Examples: Fe_2O_3; 2(Fe) + 3(O)=0, 2(Fe) + 3(-2)=0, Fe=+3 CO_2; (C) + 2(O)=0, (C) + 2(-2)=0, C=+4 Na_2O; 2(Na) + (O)=0, 2(Na) + (-2)=0, Na=+1	Examples: Na_2O_2; 2(Na) + 2(O)=0, 2(+1) + 2(O)=0, O= -1 BaO_2; (Ba) + 2(O)=0, (+2) + 2(O)=0, O= -1
hydrogen	**+1(hydrogen with a nonmetal)**	**-1(hydride = hydrogen and a metal)**
	Examples: H_2S; 2(H) + (S)=0, 2(+1) + (S)=0, S=-2	Examples: BaH_2; (Ba) + 2(H)=0, (+2) + 2(H)=0, H=-1
Group IA	**+1 (Li, Na, K, Rb, Cs, Fr)**	
	Examples: Fr_2S; 2(Fr) + (S)=0, 2(+1) + (S)=0, S=-2	Rb_3N; 3(Rb) + (N)=0, 3(+1) + (N)=0, N=-3
Group IIA	**+2 (Be, Mg, Ca, Ba, Sr, Ra)**	
	Examples: BaS; (Ba) + (S)=0, (+2) + (S)=0, S=-2	Sr_3N_2; 3(Sr) + 2(N)=0, 3(+2) + 2(N)=0, N=-3
Aluminum	**+3**	
	Examples: Al_2S_3; 2(Al) + 3(S)=0, 2(+3) + 3(S)=0, S= -2	2AlN; (Al) + (N)=0, (+3) +(N)=0, N=-3
Halogens	**-1 (a halogen with only a metal), F, Cl, Br, I**	
	Examples: SbI_3; (Sb)+3(I)=0, (Sb)+3(-1)=0, Sb=+3	$PbCl_4$; Pb+4(Cl)=0, Pb+4(Cl)=0, Pb=+4

Things to Remember

Nitrogen in a nitrate, ($\underline{N}O_3^{-1}$) ion is always +5.

Examples: $Ca(\underline{N}O_3)_2$; (Ca)+2(N)+6(O)=0, (+2)+2(N)+6(-2)=0, N=+5

$Al(\underline{N}O_3)_3$; (Al)+3(N)+9(O)=0, (+3)+3(N)+9(-2)=0, N=+5

When a metal is combined with a nonmetal, the metal is the positive element.

Sulfur in a sulfate, ($\underline{S}O_4^{-2}$) ion is always +6.

Examples:

$Ca\underline{S}O_4$; (Ca)+(S)+4(O)=0, (+2)+(S)+4(-2)=0, S=+6

$Al_2(\underline{S}O_4)_3$; (Al)+3(S)+12(O)=0, (+3)+3(S)+12(-2)=0, S=+6

An uncombined element has the oxidation number of zero (0).